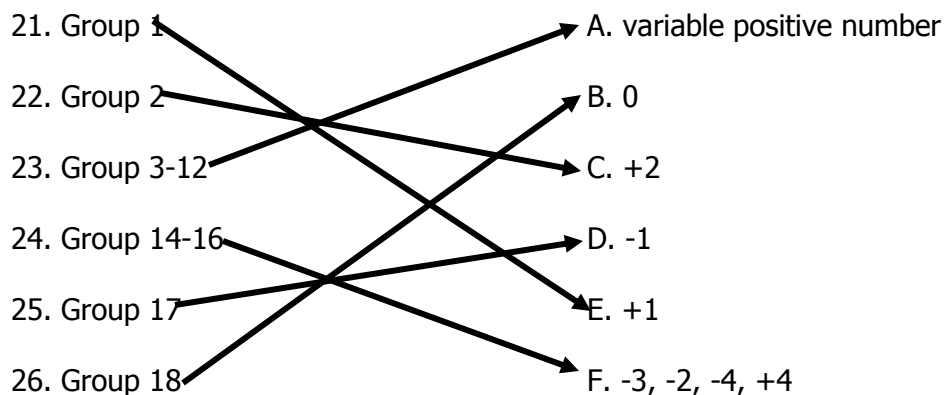


## Chemistry I: Periodic Table Review

True or False: If the statement is false, correct the statement to read true.

1. The columns on the Periodic Table are called periods. **FALSE**
2. Elements in columns will have similar chemical properties and physical properties. **TRUE**
3. The rows on the Periodic Table are called triads. **FALSE**
4. The elements in Group 1 (IA) are called alkali metals. **TRUE**
5. The elements in Group 2 (IIA) are called transition metals. **FALSE**
6. The elements in Groups 3 - 12 are called alkaline earth metals. **FALSE**
7. The elements in Group 17 (VIIA) are called halogens. **TRUE**
8. The Periodic Law states that the chemical and physical properties of the elements are a function of the atomic number. **TRUE**
9. The scientist who developed an early Periodic Table in 1869 is Moseley. **FALSE**
10. Elements found on the left side of the Periodic Table are called nonmetals. **FALSE**
11. Elements found on the right side of the Periodic Table are called nonmetals. **TRUE**
12. Those elements such as B, Si, and As that have both metallic and nonmetallic properties are called metalloids. **TRUE**
13. In the first Periodic Table, the scientist left spaces (blanks) to predict new elements. **TRUE**
14. The Noble elements can be found in Group 17. **FALSE**
15. As you go down a group on the Periodic Table, the atomic radius will decrease. **FALSE**
16. As you go across a period (L to R) on the Periodic Table, atomic radius will decrease. **TRUE**
17. As you go down a group on the Periodic Table, ionization energy will decrease. **TRUE**
18. As you go across a period (L to R) on the Periodic Table, ionization energy will decrease. **FALSE**
19. As you go down a group on the Periodic Table, electronegativity will decrease. **TRUE**
20. As you go across a period (L to R) on the Periodic Table, electronegativity will decrease. **FALSE**

Matching: Match the oxidation number with the group.



Circle the correct answer.

<b>Na</b>	Mg	Ar	member of the Alkali Metal family
Al	<b>Cl</b>	Br	gas at room temperature
O	<b>S</b>	Se	3 energy levels
O	<b>F</b>	Ne	most reactive element
Kr	I	Be	member of the Halogen family
Mg	Ca	<b>Sr</b>	largest atomic radius
<b>Ne</b>	O	F	highest ionization energy
N	<b>Mg</b>	As	always forms positive ions when bonding
N	O	<b>F</b>	smallest atomic radius
<b>K</b>	N	B	metal
<b>Xe</b>	H	Li	member of the Noble Gas family
Br	Cl	<b>F</b>	highest electronegativity value
Al	<b>Si</b>	P	metalloid
Ne	Kr	<b>Rn</b>	radioactive member of the Noble family
K	<b>Cs</b>	Rb	most reactive of the Group 1 elements
N	O	<b>F</b>	highest electronegative value
<b>N</b>	P	As	highest ionization energy